

COURSE GUIDE

2022/23

Faculty 215 - Faculty of Chemistry

Cycle Not Applicable

Degree GQUIMI20 - Bachelor's Degree in Chemistry

Year Second year

COURSE

26122 - Experiments in Physical Chemistry

Credits, ECTS: 6

COURSE DESCRIPTION

This subject introduces students to some experimental methods to determine macroscopic properties of matter in aspects concerning Chemical Thermodynamics, Chemical Kinetics and electrochemical phenomenon. In addition, previous knowledge will also be applied concerning surface properties, macromolecules and colloids.

COMPETENCIES/LEARNING RESULTS FOR THE SUBJECT

The subject is included in Physical Chemistry, within the Fundamental Module of the Degree Course and, as such, shares the transverse skills [M02CM08], [M02CM09], [M02CM10] and [M02CM11] assigned to this module. In addition, this subject develops specific skills [M02CM01]. Finally, the transverse skill of titration will also be covered [G008].

M02CM01. Understanding and handling the principles of Physical Chemistry and its influence on chemical processes.

M02CM08. Ability to select different instrumental techniques (simple or combined) to characterise chemical substances.

M02CM09. Be able to comprehensibly present phenomena and processes related to Chemistry and associated topics, both verbally and in writing.

M02CM10. Information search and selection skills in the field of Chemistry and other scientific areas, making use of the bibliography and information and communication technologies.

M02CM11. Be capable of relating Chemistry to other disciplines, and understanding its impact on today's society and the importance of the industrial chemical sector.

G008. Demonstrate teamwork and problem-solving skills in multidisciplinary contexts.

CONTENIDOS TEÓRICO-PRÁCTICOS

Chemical Thermodynamics and Thermochemistry: Determining changes in enthalpy associated with different physical-chemical processes (reaction, dissolving, etc.) using calorimetric devices.

Binary mixtures: Determining partial molar volumes of different liquid mixtures at different temperatures.

Phase equilibria: Determining vapour pressure. Phase diagrams for different solid-liquid and liquid-vapour systems.

Liquid/liquid diagrams.

Chemical equilibria: Determining equilibrium constants.

Electrochemistry: Determining potentiometry. Electrode potentials. Thermodynamic magnitudes in electrochemical batteries.

Chemical kinetics: Determining constants and reaction orders for certain reactions. Influence of temperature.

Surface and colloid phenomena: Surface tension of liquids and solutions. Surfactants: Critical micellar concentration.

Surface adsorption phenomenon.

Macromolecules. Magnitudes in solution: viscosimetry. Determining molecular weights.

TEACHING METHODS

The program is divided into two semesters and we will use traditional methodology and it will be assessed, taking into account the aforementioned skills, in the following way:

The written exam counts for 40% of the mark. Skills [M02CM01] and [M02CM08].

The remaining 60% is split between a) complete individual reports and oral presentations (30%) ([M02CM08], [M02CM09], [M02CM10] and [M02CM11]) b) teamwork and continuous assessment (%10) ([M02CM01], [M02CM08], [M02CM11] and [G008]) and c) laboratory notebook (%20) (M02CM09), [M02CM03].

The mark for each semester will count for 50 % of the final mark.

TYPES OF TEACHING

Types of teaching	M	S	GA	GL	GO	GCL	TA	TI	GCA
Hours of face-to-face teaching		5	10	45					
Horas de Actividad No Presencial del Alumno/a		7,5	15	67,5					

Legend: M: Lecture-based

S: Seminar

GA: Applied classroom-based groups

GL: Applied laboratory-based groups

GO: Applied computer-based groups

GCL: Applied clinical-based groups

TA: Workshop

TI: Industrial workshop

GCA: Applied fieldwork groups

Evaluation methods

- Continuous evaluation
- End-of-course evaluation

Evaluation tools and percentages of final mark

- Written test, open questions 40%
- Individual assignments 20%
- Teamwork assignments (problem solving, Project design) 10%
- Oral presentation of assigned tasks, Reading¿ 30%

ORDINARY EXAMINATION PERIOD: GUIDELINES AND OPTING OUT

TO BE ABLE TO MAKE THE MEDIA, IT WILL BE NECESSARY TO OBTAIN A MINIMUM OF 4/10 IN ALL AND EACH OF THE ACTIVITIES

For the student to be officially 'Withdrawn', they should drop out of the relevant activities one month before completing the classroom-based credits , notifying the teacher in writing, signed. In this case, the student will not take either the first sitting or the resit exams.

EXTRAORDINARY EXAMINATION PERIOD: GUIDELINES AND OPTING OUT

The student could be asked to complete the assessed activities that were not passed in the first sitting and/or a written exam. The marks for assessed activities that were passed in the first sitting will be passed over into the resits.

Given the nature of continuous assessment, the student will have a mark from the first sitting in May independently of whether they attend the final exam or not.

MANDATORY MATERIALS

Material necessary for laboratory practices: gown, safety glasses, laboratory notebook and scientific calculator.

BIBLIOGRAFÍA

Basic bibliography

A. M. Halpern, G. C. McBane. Experimental Physical Chemistry. A laboratory textbook, 3rd edition. W.H. Freeman, 2006.

C. W. Garland, J. W. Nibler, D. P. Shoemaker. Experiments in Physical Chemistry, 7th Edition. McGraw-Hill, 2002.

R. J. Sime. Physical Chemistry: Methods, Techniques, and Experiments. Saunders College Publishing, USA, 1990.

J. J. Ruiz Sánchez, J.M. Rodríguez Mellado, E. Muñoz Gutiérrez, J. M. Sevilla Suárez de Urbina. Curso Experimental en Química Física. Ed. Síntesis, 2003.

Detailed bibliography

D. R. Lide ed., Handbook of Chemistry and Physics, 89th Edition, CRC Press, 2008

I. R. Levine. Fisicoquímica, vols. 1 y 2. 5º ed. Ed. Mac Graw Hill, 2004.

P. Atkins, J. de Paula. Química Física. Ed. Panamericana, 2008.

R. J. Silbey, R. A. Alberty. Kimika Fisikoa. Euskal Herriko Unibertsitatea, 2006.

Journals

Journal of Physical Chemistry

Journal of Chemical Physics

Journal of Chemical Education

Web sites of interest

<http://webbook.nist.gov/chemistry>

<http://bcs.whfreeman.com/pchem8e>

<http://www.shu.ac.uk/schools/sci/chem/tutorials/>

OBSERVATIONS

On the e-gela page, students can consult the section or guidelines used to assess their oral presentations, reports, etc.

Laboratory practicals are compulsory. Students who do not attend practicals cannot pass the subject and will not be assessed in either the first sitting or resits of exams.

Any student who wishes to exercise their right to a single assessment should follow the procedure approved by the Faculty Committee on 10-10-2012.